

KAKAMEGA COUNTY JOINT EVALUATION TEST

CHEMISTRY PRACTICAL TEST-2014

CONFIDENTIAL INSTRUCTIONS

Each candidate must have the following

- 1) 50cm³ of solution **A**
- 2) 120cm³ of solution **B**
- 3) 1.0g of solid **S**
- 4) 0.5g of solid **C**
- 5) 50ml burette
- 6) 25.0ml pipette.
- 7) One 250ml beaker
- 8) One 100ml measuring cylinder
- 9) One 250ml volumetric flask
- 10) 500ml distilled water in a wash bottle
- 11) Two labels
- 12) Two 250ml conical flasks
- 13) 60 cm³ of solution **P**
- 14) 80cm³ of solution **Q**
- 15) Thermometer (0⁰C – 100⁰C)
- 16) 100ml plastic beaker
- 17) Small piece of tissue paper
- 18) Stop clock/ watch
- 19) 4 clean test tubes in a rack
- 20) One boiling tube
- 21) About 1.0g of sodium carbonate
- 22) One spatula.

Access to

- (i) Methyl orange indicator
- (ii) Acidified potassium manganate (VII) solution
- (iii) Bromine water
- (iv) Source of heat

NOTE

- (i) Solution **A** is prepared by dissolving 80cm³ of concentrated hydrochloric acid. (Conc. HCl. S.G=1.18g. cm⁻³, 36% assay) in 400cm³ of distilled water and diluting it to one litre of solution.
- (ii) Solution **B** is prepared by dissolving 8.0g of sodium hydroxide pellets in 400cm³ of distilled water and diluting it to one litre of solution.
- (iii) Solid **C** is prepared by crushing egg shell (not from boiled eggs) into powder and giving each candidate exactly 0.5g weighed accurately.
- (iv) Solution **P** – 1M- NaOH

- (v) Solution **Q** made by dissolving 60g of oxalic acid in about 600cm³ of distilled water and solution made to one litre.
- (vi) Solid **S** is oxalic acid

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