

233/3

CHEMISTRY**PAPER 3 (PRACTICAL)****CONFIDENTIAL****KAHURO/MURANG'A EAST JOINT EXAMINATION – 2016**

In addition to the common apparatus and fittings found in a Chemistry laboratory, each candidate should be provided with;

- Solid A (Six pieces of 1cm long magnesium ribbon)
- 1 boiling tube.
- Distilled water in a wash bottle.
- 1 filter funnel.
- 1 filter paper.
- 1 label.
- Solid B (1g).
- Solid W (1g).
- Universal indicator paper and a pH chart.
- About 60cm³ of solution Y 0.4M sodium carbonate solution.
- About 120cm³ of solution X 2M hydrochloric acid.
- A pipette (25ml).
- A burette.
- 10ml measuring cylinder.
- A stop watch.
- 100ml beaker.
- Stirring rod.
- Sodium hydrogen carbonate (About 1 g).

ACCESS TO:

- 2M NaOH solution.
- Ammonia solution (2M).
- Bunsen burner.
- Lead (II) nitrate solution.
- Acidified potassium manganate (VII) solution.
- Methyl orange indicator.
- Dilute nitric (V) acid.
- Sodium chloride solution.

NB:

- Solid B is a mixture of sodium sulphate and lead (II) carbonate in a ratio of 1: 1. Prepare solid B for each candidate separately.
- Solid W is maleic acid.