GATUNDU SUB COUNTY FORM FOUR 2014 EVALUATION EXAM
233/3
CHEMISTRY PAPER 3
JULY/AUGUST 2014

CONFIDENTIAL

INSTRUCTIONS TO SCHOOL

1. Only the teacher in-charge of the Chemistry practical and the school head should handle this paper.
2. Ensure that information herein does not reach the candidates either directly or indirectly.

In addition to the apparatus and fittings found in the laboratory, each candidate will require the following:

A.

1. 6g of accurately weighed solid A in a stoppered container.
2. About 100cm^3 of solution B
3. One burette
4. One pipette
5. One thermometer
6. One spatula (metallic)
7. Two boiling tubes
8. One 250ml volumetric flask
9. Three 250ml conical flasks
10. Two labels
11. 500ml distilled water
12. One clean glass rod
13. Six test tubes in a test tube rack
14. One test-tube holder
15. One watch glass
16. 250ml beaker half-filled with cold water (room temperature)
17. 6cm^3 of liquid J.
18. Solid sodium hydrogen carbonate (about 1g)

Access to:-

1. Bunsen burner
2. Phenolphthalein indicator supplied with a dropper
3. 1M potassium iodide solution supplied with a dropper.
4. 1M barium chloride solution supplied with a dropper.
5. 2M hydrochloric acid.
6. Acidified potassium dichromate (vi)
7. Acidified potassium Manganate (vii)
8. Bromine water

NB:

1. Solid A = 6g maleic acid accurately weighed
2. solid T = sodium sulphite (about 2g)
3. Dissolve 25g in 200cm^3 of 2M sulphuric (vi) acid and make to 1dm^3 mark to prepare $K_2Cr_2O_7$.
4. Liquid J = Absolute ethanol
5. Dissolve 12.5g in 400cm^3 of 2M sulphuric (vi) acid and make to 1dm^3 mark to prepare $KMnO_4$.
6. Solution B is prepared by dissolving 8g of sodium hydroxide pellets in 400cm^3 of distilled water and make it to 1dm^3 mark.

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