**TRANS MARA EAST SUB-COUNTY JOINT EVALUATION**

**EXAMS -2015**

**233/3**

**CHEMISTRY.**

**PAPER 3**

**PRACTICALS**

**CONFIDENTIAL INSTRUCTIONS TO SCHOOLS**

The information contained in this paper is to enable the head of school and teacher in charge of Chemistry to make adequate preparations for this year’s Chemistry mock practical examination. NO ONE ELSE should have access to this paper or acquire knowledge of its contents. Great care must be taken to ensure that the information herein does not reach the candidates either directly or indirectly.

The Chemistry teacher is NOT expected to perform the experiments

The apparatus required by each candidate for the Chemistry mock practical examination are set out on the next page. It is expected that the ordinary apparatus of a Chemistry laboratory will be available.

The Chemistry teacher should note that it is his/her responsibility to ensure that each apparatus acquired, for this examination agrees with specifications on the next page.

***Each candidate will require the following*:**

* One 10ml measuring cylinder
* Two boiling tubes
* One 0-25ml pipette
* One -50ml burette
* One pipette filler
* Two conical flasks
* One 50ml volumetric flask
* Six clean dry test –tubes
* One -10oC-110oC thermometer
* Distilled water in a wash bottle (500ml)
* 100ml measuring cylinder
* One 100ml beaker
* About 1.0g of solid R
* About 0.3 g solid G
* About 0.7g of solid M
* One label
* About 60cm3 of solution P

**Access** **to**

* 2M ammonia solution
* 0.5M barium chloride
* 0.5M lead(II) nitrate solution
* 2M potassium iodide
* 2M sodium hydroxide
* Acidified potassium manganaate
* 2M nitric acid
* Source of heat

**Preparations**

1. Solution L is prepared by dissolving 3.158g of potassium mangante(VII) in 400cm3 of 2M sulphuric acid and distilling with distilled water to 100cm3
2. Solution P is prepared by dissolving 39.875g of anhydrous copper (II) sulphate in about 400cm3 of distilled water and upto 1000cm3 of solution using distilled water.