

WESTLANDS JOINT EXAMINATION

Kenya Certificate of Secondary Education

CHEMISTRY

PRACTICAL

July / August 2016

CONFIDENTIAL INSTRUCTIONS

Instructions to schools;

The information contained in this paper is to enable the head of the school and the teacher in charge of chemistry to make adequate preparation for Chemistry practical examination. NO ONE ELSE should have access to this paper or acquire knowledge of its content. Great care must be taken to ensure that the information herein does NOT reach the candidates either directly or indirectly. The teacher in charge of chemistry should not perform any of the experiments in the same room as the candidates nor make the results of the experiments available to the candidates or give any other information related to the experiment to the candidates. Doing so will constitute an examination irregularity which is PUNISHABLE.

In addition to the normal laboratory fittings and apparatus, each candidate should have the following:

Requirements;

1. About 100cm³ of solution Q (0.582M) HCl.
2. 1.06g of solid R (anhydrous Na₂CO₃)
3. One burette 0 - 50cm³
4. One pipette 25cm³
5. One thermometer (-10 -110°C)
6. Two conical flasks of 250ml
7. 100ml plastic beaker
8. Stop watch
9. Distilled water
10. 2.85g solid S (NH₄)₂CO₃ weighed ACCURATELY and stoppered in a vial.)
11. About 0.3g of solid T (Al₂(SO₄)₃)
12. About 0.5g of Solid U Maleic acid
13. 2 boiling tubes; 7 test-tubes
14. 250 ml volumetric flask
15. 50cm³ measuring cylinder.
16. Red and blue litmus papers.
17. Metallic spatula
18. 3 labels

ACCESS TO

1. Methyl orange indicator

2. Phenolphthalein
3. 2M aqueous ammonia
4. 1 M Barium chloride solution
5. 1M Lead (II) nitrate solution.
6. Bunsen burner
7. 0.2M acidified potassium manganate (VII) solution
8. pH paper and pH chart (full range)
9. 3 labels