**CHEMISTRY 233/3 PRACTICAL**

**CONFIDENTIAL TO SCHOOLS**

 Each candidate will require

1. Solution A 100cm3
2. Solution B 150 cm3
3. Solution C 100cm3
4. Burette
5. 25ml pipette
6. 2 conical flasks
7. Retort stand
8. Filter funnel
9. Pipette filler
10. 100ml of distilled water
11. Thermometer
12. 1 Spatula – full of solid D
13. About 10ml liquid E
14. 4g solid F ( weighed exactly)
15. 2 boiling tubes
16. Six test tubes in a rack
17. Test tube holder
18. Metallic spatula

 **ACCESS TO**

* 1M NaOH
* 1M NH4OH
* 0.1M NaCl
Acidified K2Cr2O7
* KMnO4- use some amount of solution B
* Source of heating
* **NOTES**
1. **Solid A**

 Dissolve 7.0g Ferrous Sulphate (FeSO4.7H2O) in 50ml of 1MH2SO4, dilute to 1dm3 with water. (Should be prepared in the morning of the exam day)

1. **Solution B**

 Dissolve 0.8g of KMnO4 in 50cm3 of 1MH2SO4. Dilute to 1 dm3 with water.

1. **Solution C**

 Measure 3cm3 of 20vol. H2O2

 Dissolve in 1dm3 of solution.

1. **1M H2SO4**

 Measure 55cm3 of conc. H2SO4 add to about 200cm3 of water, stir, dilute to 1 dm3

1. **Solid D**

 Aluminium Nitrate

1. **Solid F**

 Potassium Chlorate (KClO3)

1. **Acidified K2Cr2O7**

 Dissolve 0.3g of K2Cr2O7 in 50cm3 of 1MH2SO4. Dilute to 1 dm3 with water.

1. **Liquid E**

 Ethanol.