

**233/3**

**CHEMISTRY**

**(PRACTICAL)**

**PAPER 3**

**CONFIDENTIAL**

Per Student

1. Solution A (100ml)
2. Solution B (100ml)
3. Phenolphthalein indicator
4. 3 conical flasks
5. Funnel
6. Burette
7. Pipette
8. Clamp
9. Stand
10. CBI (g) – NaHCO3(s)
11. Clean spatula
12. Test- tubes (5)
13. Litmus papers ( 2 blue and 2 red)
14. Distilled water
15. Solid Q – 1g (NH4)2 SO4.FeSO4. 6H2O and NaCl (ration 1:1)
16. 1 boiling tube

**Access to;**

1. 2M ammonia solution
2. 2M Sodium hydroxide solution
3. Source of heat
4. Silver nitrate solution (0.05M)
5. Dilute nitric acid (0.1M)
6. Dilute hydrochloric acid (0.1M)
7. Dilute Barium nitrate solution (0.1M)
8. Conc. Nitric acid in dropper bottles
9. White tile
10. Test tube holder

* Solution A is prepared by dissolving 6.3g of H2C2O4. 2H2O in 400cm3 of water and topped upto one litre of solution.
* Solution B is prepared by dissolving 4g of Sodium hydroxide in 400cm3 of water and topped upto one litre of solution.