**CHEMISTRY**

**PAPER 3**

**CONFIDENTIAL**

**Each candidate should be supplied with the following;**

1. Clamp and stand
2. Filter funnel
3. White tile
4. 100 ml of 0.21M glucose solution Labelled Solution V
5. 10 ml of 0.02M of potassium manganate (VII) Labelled solution W.
6. 250 ml of 1.0M aqueous sulphuric (VI) acid
7. Thermometer
8. About 250ml distilled water
9. Stop watch
10. 1 250ml beaker
11. 10ml measuring cylinder
12. 100ml measuring cylinder
13. 250cm3 volumetric flask
14. 2 conical flasks
15. 25cm3 pipette
16. 50cm3 burrette
17. Metalic spatula
18. 6 clean test tubes
19. One boiling tube
20. Glass rod
21. Metallic spatula
22. Labeling paper
23. Pipette filler
24. 500cm3of distilled water in a wash bottle
25. About 70cm3 of solution K
26. About 100cm3 of solution M
27. About 0.5g of solid P
28. About 0.5g of solid F
29. Phenolphthalein indicator

**Each student should have access to the following;**

1. 2M Ba{NO3}2
2. 2M HNO3
3. 2M NaOH
4. Acidified potassium manganate vii

{Note: all the above solutions should be supplied with a clean dropper each]

1. Means of heating

**SOLUTION PREPARATION**

1. Solution K is 1.25M HCl
2. Solution M is NaOH containing 6.4g of sodium hydroxide in1 litre of the solution
3. Solid P. sodium sulphite
4. Solid F is maleic acid
5. Acidified potassium manganate VII is prepared by dissolving 3.2g of potassium manganate VII in 200cm3 of 2M sulphuric VI acid in one litre volumetric flask, the adding distilled water to the mark
6. Solution V 0.21M glucose solution
7. Solution W 0.02M KMnO4